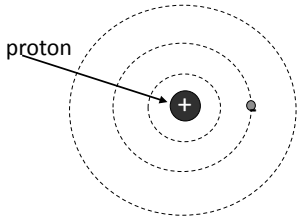
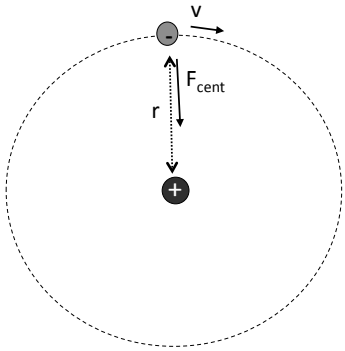
Bohr model of the atom
(not quite right)

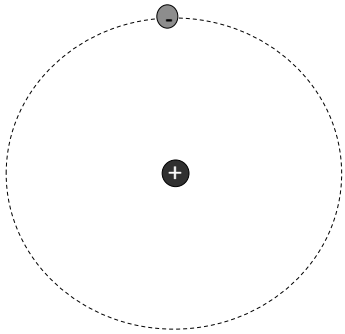


Basic connections between r, v, and energy!



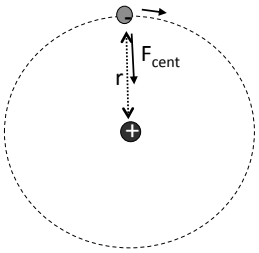
$F = ma = F_{cent} = ?$

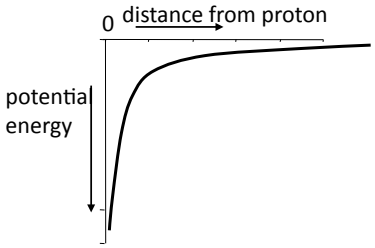
A) $-mvr$
B) $-mv^2/r$
C) $-v^2/r^2$
D) I don't remember learning anything related to this



The potential energy of the electron is:

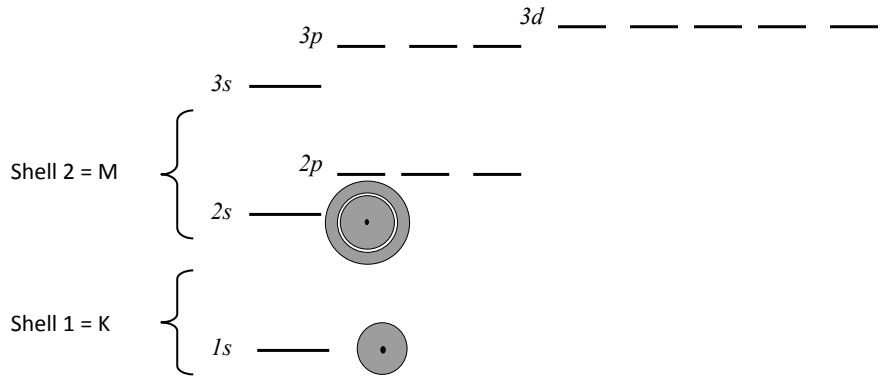
A) $-ke^2/r$
B) $-ke^2/r^2$
C) $-ke^2r$
D) I don't remember



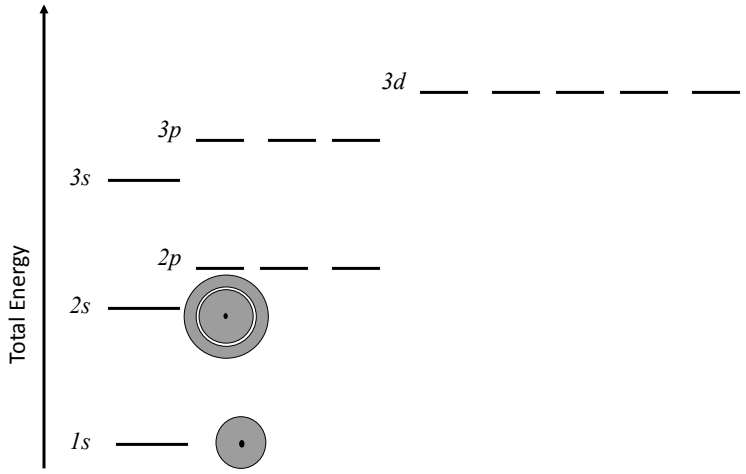


An electron in the Bohr model of hydrogen atom:

A) is always at one particular distance from the nucleus
B) can be at any distance from the nucleus.
C) is at certain distances from the nucleus corresponding to energy levels it can be in.
D) must always go into the center where potential energy is lowest.



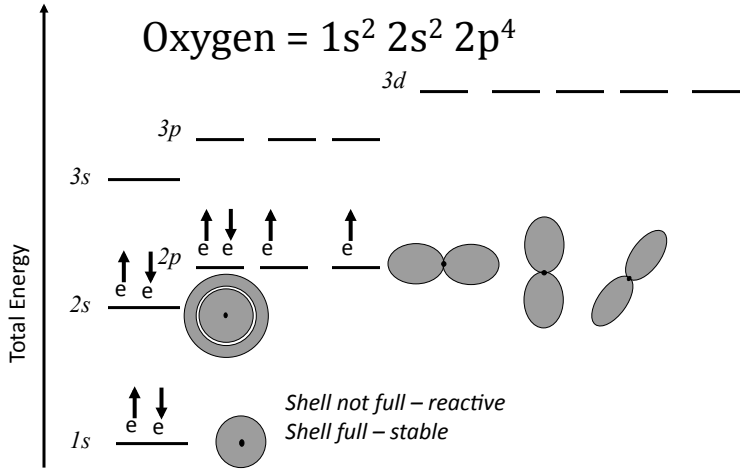
How do the electrons fit into the available orbitals?



A brief review of chemistry

Filling orbitals … lowest to highest energy, 2 e's per orbital

O
N
C
B
Be
Li
He
H



What is the electronic configuration of ^{12}Mg?

A) $1s^2\ 2s^2\ 3s^2\ 4s^2\ 5s^2\ 6s^2$

B) $1s^2\ 2s^2\ 2p^2\ 3s^2\ 3p^2\ 3d^2$

C) $1s^2\ 2s^2\ 2p^6\ 3s^2$

D) [Ne] $3s^2$